

Complete the last column using T/F to indicate whether the statement is true or false.

Use the space below each false statement to rewrite it.

		T/F
1	Collision theory states that particles must collide with enough energy in order to react.	
2	The minimum amount of energy required for a successful collision is called the starter energy.	
3	During a reaction a catalyst gets used up and has to be replaced each time.	
4	Increasing the pressure of gases in a reaction will increase the rate of reaction.	
5	Increasing the temperature of reactants increases the number of particles in the reaction.	
6	Different catalysts are needed for different reactions.	
7	There are five main factors that affect the rate of reaction.	
8	Measuring the gas produced, the mass lost or the light transmitted are all ways of measuring the rate of reaction.	
9	For every 10°C temperature increase, the rate of reaction triples.	
10	Catalysts are usually cheap transition metals.	

Answers

		T/F
1	Collision theory states that particles must collide with enough energy in order to react.	T
2	The minimum amount of energy required for a successful collision is called the starter energy. It is called the activation energy.	F
3	During a reaction a catalyst gets used up and has to be replaced each time. The catalyst in a reaction does not get used up and can be used over and over again.	F
4	Increasing the pressure of gases in a reaction will increase the rate of reaction.	T
5	Increasing the temperature of reactants increases the number of particles in the reaction. Increasing the temperature increases the speed of the particles. Moving faster means there are more successful collisions.	F
6	Different catalysts are needed for different reactions.	T
7	There are five main factors that affect the rate of reaction. There are FOUR main things that affect the rate of reaction: temperature, surface area, catalysts and concentration (pressure in gases is the same as concentration in liquids).	F
8	Measuring the gas produced, the mass lost or the light transmitted are all ways of measuring the rate of reaction.	T
9	For every 10°C temperature increase, the rate of reaction triples. For every 10°C increase, the rate of reaction roughly doubles.	F
10	Catalysts are usually cheap transition metals. Catalysts are often very expensive because they are made of precious metals.	F